Checking Why a Melt and a Solution Conduct

Scoring guide · one credit per question · review written work against each rubric

Checking Why a Melt and a Solution Conduct

Question 1

Choice (1)

Both substances give a conducting water mixture, so that result is the same for P and Q.

Question 2

In the melt of P the existing ions are free to move and carry charge. The pure melt of Q contains neutral molecules and does not have those mobile ions in this model.

Question 3

Q is molecular as a pure solid. Its molecules react with water to make mobile ions, so a conducting solution does not establish that the pure solid originally consisted of ions.

Question 4

It does not conduct under the model. Dissolving Q in that liquid leaves neutral molecules and does not supply mobile charged particles.