Electronegativity Trends Worksheet

Scoring guide · one credit per question · review written work against each rubric

Electronegativity Trends

Question 1

Choice (1)

Electronegativity refers to attraction for electrons shared in a bond.

Question 2

Chlorine is farther right in the period. Its greater effective nuclear attraction pulls shared electrons more strongly.

Question 3

Choice (2)

Chlorine is above bromine in Group 17. Atoms higher in that group generally attract shared electrons more strongly.

Question 4

Outer electrons are farther from the nucleus and are shielded by more inner electrons. The weaker attraction generally lowers electronegativity down the group.

Reading a Periodic-Table Excerpt

Question 5

Choice (3)

Chlorine has the largest supplied value, 3.16.

Question 6

1.61 − 0.93 = 0.68.

Question 7

Choice (4)

Cl and Br are both in Group 17 and are in different periods.

Question 8

For example, Na in Group 1 has a value of 0.93 and Cl in Group 17 has 3.16. The value increases between these two examples, contradicting the claim. Any valid pair of Period 3 rows earns credit.