# Separating Made Water From Incoming Water

Scoring guide · one credit per question · review written work against each rubric

### Separating Made Water From Incoming Water

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### Question 1

Each reaction consumes two moles of its fuel per mole of $O_2$. Regardless of the mixture, 0.050 mol $O_2$ corresponds to 2 × 0.050 = 0.100 mol of $H_2$ and CO combined.

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### Question 2

Choice (2)

The second equation gives one mole of $CO_2$ per mole of CO consumed, so 0.040 mol CO was used.

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### Question 3

The amount of $H_2$ consumed is 0.100 − 0.040 = 0.060 mol. It makes 0.060 mol $H_2O$. The difference, 0.080 − 0.060 = 0.020 mol, was carried in as water.

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### Question 4

The water reading drops to 0.060 mol, the amount made by the reaction. The oxygen and carbon dioxide readings stay the same because the amounts of hydrogen and carbon monoxide are unchanged.